

1 Patent Application

of

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for

6 **SELF-MOISTURIZING PROTON EXCHANGE MEMBRANE, MEMBRANE-
ELECTRODE ASSEMBLY AND FUEL CELL**

FIELD OF THE INVENTION

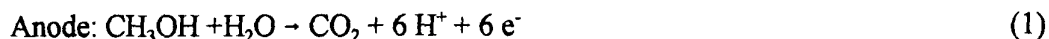
11 This invention relates generally to a polymer electrolyte membrane (PEM) for use in a
fuel cell, and more particularly to a self-moisturizing PEM, membrane/electrode assembly and
fuel cell.

BACKGROUND OF THE INVENTION

16 A fuel cell converts chemical energy into electrical energy and some thermal energy by
means of a chemical reaction between hydrogen-containing fuel and oxygen. As compared to
other energy sources, fuel cells provide advantages that include low pollution, high efficiency,
high energy density and simple fuel recharge. Fuel cells can be used in electrochemical engines,
portable power supplies for various microelectronic and communication devices, standby power
supply facilities, power generating systems, etc. Further, fuel cells utilize renewable resources
and provide an alternative to burning fossil fuels to generate power.

21 The chemical reaction of a fuel cell requires the presence of an electrolyte, electrodes
and catalysts. Based on the electrolyte type, the fuel cell is classified as alkaline fuel cell (AFC),
phosphoric acid fuel cell (PAFC), molten carbonate fuel cell (MCFC), solid oxide fuel cell
(SOFC), and proton exchange membrane or polymer electrolyte membrane fuel cell (PEM-FC).
26 The PEM-type fuel cell has been a topic of highly active R&D efforts during the past decade. A
PEM-FC uses hydrogen or hydrogen-rich reformed gases as the fuel while a direct-methanol fuel
cell (DM-FC) uses methanol solution. A DM-FC also contains a PEM interposed between the
two electrodes. PEM-FC and DM-FC are collectively referred to as the PEM-type fuel cell.

1 In a DM-FC, a mixture of methanol and water is circulated past an anode that is
separated from a cathode by a membrane that is selectively permeable to protons. The following
chemical reaction takes place at the anode.



6 The protons generated at the anode pass through the membrane to the cathode side of the fuel
cell. The electrons generated at the anode travel to the cathode side of the fuel cell by passing
through an external load that connects the anode and cathode. Air or an alternative oxygen
source is present at the cathode where the electro-reduction of oxygen occurs resulting in the
following chemical reaction:



11 A PEM-type fuel cell is typically composed of a seven-layered structure, including a
central polymer electrolyte membrane for proton transport, two electro-catalyst layers on the two
opposite sides of the electrolyte membrane in which the chemical reactions occur, two gas
diffusion electrodes (GDEs) or backing layers stacked on the corresponding electro-catalyst
16 layers (each GDE comprising porous carbon paper or cloth through which reactants and reaction
products diffuse in and out of the cell), and two flow field plates stacked on the GDEs. The flow
field plates are made of carbon, metal, or composite graphite fiber plates which also serve as
current collectors. Gas-guiding channels are defined on a GDE facing a flow field plate, or on a
flow field plate surface facing a GDE. Reactants and reaction products (e.g., water) are guided
21 to flow into or out of the cell through the flow field plates. The configuration mentioned above
forms a basic fuel cell unit. Conventionally, a fuel cell stack comprises a number of basic fuel
cell units that are electrically connected in series to provide a desired output voltage. If desired,
cooling plates and humidifying plates may be added to assist in the operation of a fuel cell stack.

26 Two or three of the above-described seven (7) layers may be integrated into a compact
assembly, e.g., the membrane-electrode assembly (MEA). The MEA typically includes a
selectively permeable polymer electrolyte membrane bonded between two electrodes (an anode
and a cathode). Typically, the anode and the cathode is each bonded to an electro-catalyst layer,
as described above, or each contains an electro-catalyst phase dispersed therein. Commonly

1 used electro-catalysts include noble metals (e.g., Pt), rare-earth metals (e.g., Ru), and their alloys. Known processes for fabricating high performance MEAs involve painting, spraying, screen-printing and hot-bonding catalyst layers onto the electrolyte membrane and/or the electrodes.

6 Hydrogen transport through the PEM layer in a PEM-type fuel cell requires the presence of water molecules within the membrane such as poly(perfluoro sulfonic acid), its derivative, copolymer, or mixture. Consequently, it is critical to maintain adequate membrane hydration in order for the fuel cell to function properly. In addition to maintaining adequate ionic conductivity and proton transport, uniform membrane hydration serves to prevent localized
11 drying, or hot spots, that could result from higher localized resistance. In general, dehydration may impede performance, increase resistive power losses and degrade the structure of the membrane.

16 In conventional fuel cells, membrane hydration is achieved by humidifying the fuel (e.g. hydrogen gas) and oxidant gases (oxygen or air) prior to their introduction into the fuel cell. One commonly used method for pre-humidifying fuel cell gas streams is to employ membrane-based humidifiers (e.g., Reid, U.S. Pat. No.6,403,249, June 11, 2002). In these situations, reactant moisture is added by flowing the respective gases on one side of a water vapor exchange membrane while flowing deionized water on the opposite side of the membrane. Water is
21 transported across the membrane to humidify the fuel and oxidant gases. Another known technique for pre-humidifying the reactant gas streams entails exposing the gases directly to water in an evaporation chamber to permit the gas to absorb evaporated water. Alternatively, humidification may be achieved by directly injecting or aspirating water into the respective gas streams before introducing them into the fuel cell.

26 Generally, pre-humidification is undesirable because it requires auxiliary fuel cell components, increasing the relative complexity of fuel cell systems. For instance, pre-humidification generally requires dedicated components for storing and transporting water. Auxiliary water storage and transport components reduce operating efficiency and add to the

1 overall weight and cost of the system. Additional weight is an undesirable feature for a fuel cell
if the cell is to be used in a portable microelectronic device such as a mobile phone or a personal
data assistant (PDA). Additional components may also present system reliability issues. For
example, where fuel cells are operated in sub-freezing conditions, water solidification can result
in the weakening of mechanical components.

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Wynne, et al. (U.S. Pat. No. 6,207,312, March 27, 2001) disclosed a self-humidifying
fuel cell that made use of the reaction product (water) as a source of PEM moisture, avoiding the
use of auxiliary components. However, this fuel cell requires the design and construction of
complex flow field channels in the gas diffusion electrodes or the flow field plates.

11 SUMMARY OF THE INVENTION

It is an object of this invention to provide a self-moisturizing polymer electrolyte
membrane (PEM), a membrane-electrode assembly (MEA) and a fuel cell, which are capable of
maintaining membrane hydration during the storage and operation of the fuel cell.

16 It is another object of this invention to provide a PEM fuel cell in which membrane
hydration is maintained using water provided originally in the PEM material or water as a
byproduct of electrochemical reactions within the fuel cell.

It is a further object of this invention to provide a PEM fuel cell which does not require
components dedicated to pre-humidification of reactant gas streams.

21 These and other objects of the invention are achieved by the PEM composition, the MEA
and the fuel cell of the present invention. A self-moisturizing PEM-type fuel cell has a
membrane electrode assembly (MEA), comprising an ion-exchange membrane interposed
between catalyzed anode and cathode electrodes. The MEA is interposed between a pair of gas
26 diffusion backings, and the resulting structure is interposed between fuel and oxidant flow field
plates. The fuel flow field plate has a fuel stream inlet, a fuel stream outlet, and means for
flowing the fuel stream therebetween. The oxidant flow field plate has an oxidant stream inlet,
an oxidant stream outlet, and means for flowing the oxidant stream therebetween.

1 In a preferred embodiment, the PEM composition comprises a proton-conducting
polymer and a deliquescent material dispersed in this polymer. Suitable deliquescent materials
include, but are not limited to, calcium chloride, calcium bromide, potassium biphosphate,
potassium acetate and combinations thereof. A deliquescent material absorbs an essentially
constant amount of moisture to keep the proton mobile in the PEM structure. In another
6 preferred embodiment, a deliquescent material is strategically positioned at some points along
the fuel and/or oxidant flow paths (each from the flow field surface, through the gas diffusion
electrode or backing plate, further through the electro-catalyst layer, to the PEM layer). The
presence of a deliquescent material acts to maintain a dynamic equilibrium of water vapor
between the deliquescent material and the PEM layer.

11 It may be noted that Nitzan discloses a "paper battery" or "open cell" (U.S. Pat. Nos.
5,652,043 (July 29, 1997); 5,811,204 (Sept.22, 1998); and 5,897,522 (April 27, 1999)) that
comprises a first layer of insoluble negative pole, a second layer of insoluble positive pole and a
third layer of aqueous electrolyte, the third layer being disposed between the first and second
16 layers and including (a) a deliquescent material for keeping the open cell wet at all times, (b) an
electro-active soluble material for obtaining required ionic conductivity, and (c) a water-soluble
polymer for obtaining a required viscosity for adhering the first and second layers to the third
layer. Nitzan's approach is fundamentally different and distinct from our invention in several
ways: (1) In Nitzan's invention, the electrolyte is directly open to the atmospheric air through
21 the four side surfaces. The water vapor in the atmosphere is up-taken by the deliquescent
material dissolved in the electrolyte to keep the aqueous electrolyte hydrated. In contrast, the
PEM layer in our fuel cell is the innermost layer and is not directly open to the outside air. The
PEM is a polymer solid electrolyte, not a liquid electrolyte. The moisture needed to keep the
protons mobile could come from the water added during the fuel cell manufacturing process,
26 from the fuel or oxidant streams (if they naturally contain some moisture), and/or from the
reaction. (2) Since there is no external casing in Nitzan's battery configuration, a water-soluble
polymer has to be dissolved in the electrolyte to provide bonding between the electrolyte layer
and the two respective electrodes. By contrast, water does not dissolve the electrolyte in our fuel
cell since PEM is not water soluble. (3) Nitzan's open cell is for a battery, not a fuel cell.

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BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1 Schematic of a PEM fuel cell assembly in accordance with the present invention.

DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENTS

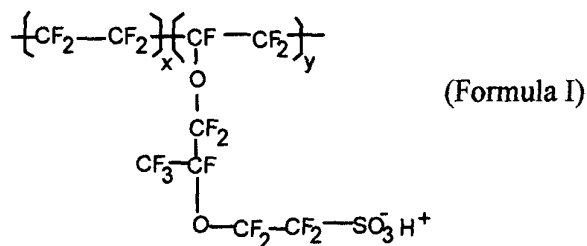
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The present invention provides a self-moisturizing polymer electrolyte membrane (PEM) composition, a membrane-electrode assembly, and a fuel cell. The PEM composition comprises (a) a proton-conducting polymer containing a detachable hydrogen ion and a counter-ion bonded to the polymer; and (b) a deliquescent material for keeping the membrane wet, thereby helping to detach the hydrogen ion to facilitate proton transport in the membrane.

11

In one preferred embodiment, the PEM is an ion exchange membranes having sulfonic acid groups. These materials hydrate when immersed in water, with hydrogen ion H^+ detached from sulfonic ion, SO_3^- . The general structure of the sulfonic acid membranes that have received extensive attention for use in fuel cells and are sold under the trade name Nafion[®] by E. I. du Pont Company is as follows:

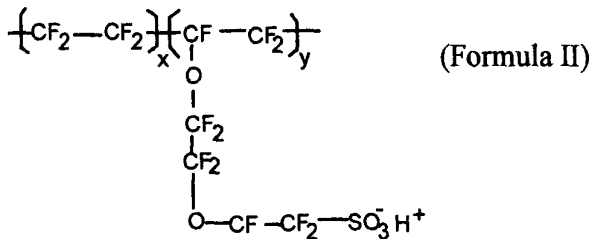
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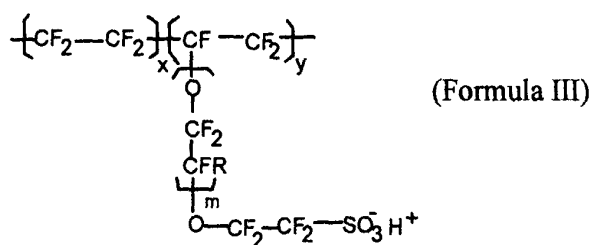
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where x and y are integers selected from 1 to 100,000, preferably from 1 to 20,000, most preferably from 100 to 10,000. A similar polymer that is also suitable for use as a PEM is given as:

26



Sulfonic acid polymers having a shorter chain between the pendant functional group (side group) and the main polymer backbone absorb less water at a given concentration of functional group in the polymer than do polymers having the general structure as shown by Formula I and II. The concentration of functional group in the dry polymer is expressed as an equivalent weight. Equivalent weight is defined, and conveniently determined by standard acid-base titration, as the formula weight of the polymer having the functional group in the acid form required to neutralize one equivalent of base. In a more general form, this group of proton-conducting polymers may be represented by the formula:



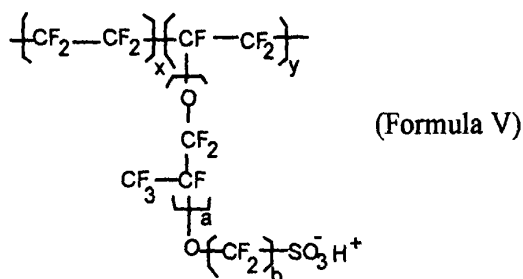
where x and y are integers selected from 1 to 100,000, m is an integer selected from 0 to 10 and R is a functional group selected from the group consisting of H, F, Cl, Br, I, and CH₃.

Another class of PEM polymers suitable for use in the fuel cell is characterized by a structure having a substantially fluorinated backbone which has recurring pendant groups attached thereto and represented by the general formula:



where a = 0-3, b = 0-3, a+b = at least 1, R_f and R_f' are independently selected from the group consisting of a halogen and a substantially fluorinated alkyl group having one or more carbon atoms.

Still another group of sulphonic acid polymers suitable for the present fuel cell PEM applications comprises a repeating unit represented by the following formula:



where a is 0, 1, or 2, b is 2 or 3, x and y are positive integer numbers and x/y ratio is of 10 or less. This polymer was found by Noaki, et al. (U.S. Pat. No. 5,449,697, Sept. 12, 1995) to be a good fuel cell PEM material. We have found that the addition of a deliquescent material helps to maintain an essentially constant moisture level in the sulfonic acid polymer-based PEM layer of a fuel cell.

The above polymers have a detachable hydrogen ion (proton) that is weakly attached to a counter-ion (e.g., SO_3^-), which is covalently bonded to a pendant group of the polymer. While the general structures shown above are representative of several groups of polymers of the present invention, they are not intended to limit the scope of the present invention. It would become obvious to those skilled in the art, from the relationships presented herein that other sulfonic acid functional polymers having pendant chains, sterically hindered sulfonate groups or the like would absorb some water and conduct protons. For instance, the derivatives and copolymers of the aforementioned sulfonic acid polymers, alone or in combination with other polymers to form polymer blends, may also be used as PEM in the invented fuel cell.

A deliquescent material refers to a hygroscopic agent, a desiccant composition, a water absorbent, or a moisture exchange element or compound. These materials are able to absorb the moisture from the atmospheric air and keep a material or structure moistened. In an enclosed or partially closed container, these materials can help maintain a dynamic moisture equilibrium between these materials and a target material inside the container.

Atmospheric water vapor can be utilized for useful purposes through the use of many compounds which are extremely hygroscopic in nature. A hygroscopic material can absorb

1 water out of the atmosphere causing the material or its mixture become deliquescent. Indeed,
there are some solid compounds which are so hygroscopic that they will literally create their
own self solutions, and continue to dilute themselves even when a solution is formed because the
solutions of these compounds are in turn hygroscopic.

6 The membrane composition of this invention comprises a combination of a strongly
hygroscopic agent, and a proton conducting polymer. The hygroscopic agents are preferably
those which will form a pasty mix by absorbing atmospheric water vapor. A variety of metal
halides such as aluminum chloride, magnesium chloride, calcium chloride, zinc chloride, or iron
chloride are very useful as the hygroscopic material. In addition, metal nitrates, such as zinc
11 nitrate, magnesium nitrate, and iron nitrate, can also be used as the hygroscopic agent.

There are several organic ions which form deliquescent salts which are also sufficiently
hygroscopic to serve in this invention. These include sodium formate, sodium ethyl sulfate and
magnesium acetate. In addition, purely organic materials, such as polyethylene glycol or
16 polyvinyl pyrrolidone are useful in absorbing water from the atmosphere.

In one preferred embodiment, the deliquescent material is uniformly dispersed in the
PEM polymer. This may be achieved by dissolving or dispersing both the PEM polymer and the
deliquescent material in a common solvent or dispersing medium to form a solution or
21 suspension. The solution or suspension is then made into a layer of modified PEM by spin-
casting, printing, spraying, or any other polymer processing technique that involves eventually
removing the common solvent or medium. Typically, the volume fraction of the deliquescent
material in the PEM composition (deliquescent material + PEM polymer) may vary from 0.5%
to 50%, but preferably from 5% to 25%. When the volume fraction of the deliquescent material
26 exceeds 30-40%, the over-all proton conductivity in the PEM composition begins to decrease
significantly.

The deliquescent material by being hygroscopic maintains the membrane moisturized at
all times. The level of moisture within the PEM may vary depending on the deliquescent

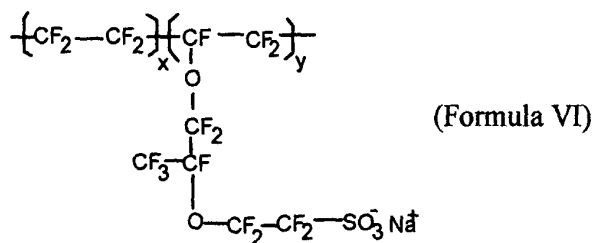
material selected, its concentration and air humidity. Suitable deliquescent materials include, but are not limited to, zinc chloride, calcium chloride, magnesium chloride, lithium chloride, calcium bromide, potassium biphosphate, potassium acetate, phosphorous oxide, ammonium acetate, sodium acetate, sodium silicate, potassium silicate, magnesium sulfate, aluminum oxide, calcium oxide, silicon oxide, zeolite, barium oxide, cobalt chloride, bentonite, montmorillonite clay, silica gel, molecular sieve, monohydric compounds, polyhydric compounds, metal nitrate salt, sodium ethyl sulfate organic salt, polyethylene glycol, and combinations thereof.

EXAMPLE 1

A solution containing 240 mg of poly (perfluoro sulfonic acid) (PPSA) and 120 mg of zinc chloride (a deliquescent material) in 3.0 ml of ethanol was prepared. This solution was then cast onto a piece of glass with the solvent evaporated in a chemical fume hood to form a layer of PPSA-ZnCl₂ mixture (sample 1B). A baseline sample (sample 1A) containing only PPSA, without any zinc chloride, was prepared by following a similar procedure. Both the mixture sample and the PPSA-only sample were placed in a low-humidity oven maintained at 80°C (a typical fuel cell operating temperature) for two weeks. The mixture (sample 1B or the PEM composition) appeared to maintain a good level of moisture, but the baseline sample (sample 1A) was dry and rigid.

EXAMPLE 2

A perfluoro sodium sulfonate type ion exchange polymer material in a powder form having an equivalent weight of 1080 g/eq represented by the following formula (VI) was prepared:

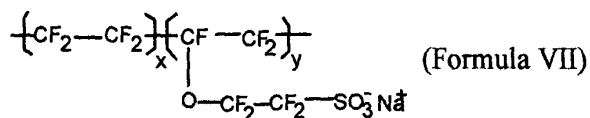


wherein x/y ratio is 6.36. The prepared ion exchange membrane was dipped in a swelling treatment liquid (ethylene glycol) at a constant temperature of 130° for 3 hours. Next, the membrane was dipped in 0.1 mol/l sodium hydroxide at 90°C for 12 hours, dipped in a 1 mol/l sulfuric acid solution at 60°C for 12 hours, and then was boiled in a mixture of water and ethanol for 2 hours. The resulting sample 2A has a structure represented by Formula I, with Na⁺ being replaced by H⁺. The above procedure is an ion exchange treatment.

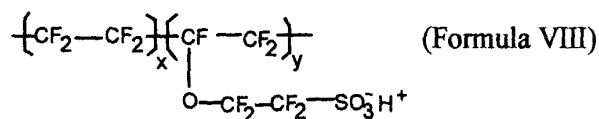
Sample 2B was prepared in a similar manner as sample 2A, with the exception that a certain amount of calcium chloride was added to the water-ethanol mixture to obtain a slurry containing the polymer, the deliquescent material (CaCl₂), and the liquid mixture. The slurry was cast onto a glass substrate to obtain a layer of PEM composition. Both sample 2A and sample 2B were placed in a low-humidity oven maintained at 80°C for two weeks. Sample 2B (PEM composition) appeared to maintain a much higher level of moisture as compared to sample 2A.

EXAMPLE 3

Sample 3A and 3B were prepared by using procedures similar to those for samples 2A and 2B, respectively, with the exception that the starting material was represented by formula VII:



and the resulting polymer after the ion exchange treatment was represented by formula VIII:



This is a special case of Formula V with a = 0 and b = 2. Again, with the presence of a deliquescent material, sample 3B maintains moisture much more effectively than does sample 3A.

EXAMPLE 4

The catalyst material (a powder of Pt dispersed on carbon) is first prepared in liquid ink form by thoroughly mixing together appropriate amounts of catalyst and a solution of the membrane composition (95% Nafion® from du Pont and 5% potassium biphosphate) in alcohol. A layer of catalyst ink was then painted onto a dry, solid piece of the same membrane composition (95% Nafion® and 5% potassium biphosphate). The wet catalyst layer and the membrane were heated until the catalyst layer was dry. The membrane was then turned over and the procedure was repeated on the other side. Catalyst layers are now on both sides of the membrane. The dry membrane/electrode assembly (MEA) was next re-hydrated by immersing in lightly boiling dilute acid solution to ensure that the membrane is in the H⁺ form needed for proton conduction. The MEA was then thoroughly rinsed in distilled water (sample 4B). A similar procedure was followed to prepare a baseline MEA comprising a layer of pure Nafion® (sample 4A).

EXAMPLE 5

For each of samples 4A and 4B, the MEA was inserted into the hardware of the fuel cell, including backing layers, flow fields, and current collectors. A porous carbon paper of approximately 100 µm thick, treated with Teflon®, was used as a backing layer. One backing layer was attached to one side of an MEA and another backing layer attached to the other side. One electronically conducting plate was then pressed against the outer surface of each backing layer. The two plates serve the dual role of flow field and current collector for both cathode and anode. The side of a plate next to the backing layer contains channels machined into the plate. The channels are used to carry the reactant gas from the point at which it enters the fuel cell to the point at which the gas exits. The two fuel cells containing 4A and 4B as an MEA are referred to as samples 5A and 5B, respectively. The two fuel cells were evaluated by allowing the two to operate in an environmental chamber where temperature and humidity were programmed to fluctuate with time. It was found that sample 5B (containing a deliquescent material) provided a relatively constant current level for a longer period of time as compared with sample 5A. Sample 5A fluctuated significantly with the humidity level, providing a lower current value when humidity was low.

1 In addition to the above five examples, more fuel cells were prepared to determine if
incorporation of a deliquescent material in different parts of a fuel cell (than the membrane
electrode) would have any effect on the performance of a fuel cell. We were surprised to
observe that a more steady, constant-current output was achieved with fuel cells that contain a
thin layer of deliquescent material coated onto the surface of an electrode, a backing layer, or
6 gas flow channels of a field plate/current collector. It appeared that the presence of a proper
amount of deliquescent materials somewhere along the flow path of a reactant gas (hydrogen or
air) helped to maintain a dynamic equilibrium of moisture inside a fuel cell. When the ambient
humidity was high, the deliquescent material became slightly more wet by absorbing more water
vapor. When the ambient humidity was low, the deliquescent material served to capture the
11 reaction product (water) and released certain amount of water back to the membrane to help
maintain the proton conductivity. It appeared more advantageous to position the deliquescent
material on the cathode side than the anode side. It also worked well when both sides contained
a deliquescent material.

16 Thus, another preferred embodiment of the present invention is a fuel cell (FIG.1)
comprising a central polymer electrolyte membrane **28** with two primary surfaces, two separate
electrode/catalyst layers **16,26** each with an inner layer attached to or integral with one of the
two primary surfaces, two backing layers **14,24** attached to the outer surfaces of the
electrode/catalyst layers, and two respective flow field/current collector plates **10,20** pressed
21 against the outer surfaces of the backing layers **14,24**. Inside the anode current collector/flow
field plate **10** are built-in channels **12** for hydrogen fuel transport. Similarly, the cathode current
collector/flow field plate **20** has flow channels **22** to admit air or oxygen and to exit water and
un-used air. The polymer electrolyte membrane **28**, the anode catalyst layer **16** and the cathode
catalyst layer **26** together constitute a membrane/electrode assembly (MEA). In this fuel cell, at
26 least one of the seven (7) layers comprises a desired amount of a deliquescent material, which
serves to self-moisturize the membrane layer. Preferably, more than one layer of the seven-layer
fuel cell structure comprises a deliquescent material. Further preferably, the deliquescent
material is coated on the surface of gas-guiding channels of a flow field plate-current collector,
or coated on a surface (either facing the flow field plate or the MEA electrode) of a GDE or

1 backing layer.

While the invention has been described with respect to a limited number of embodiments, it will be appreciated that many variations, modifications and other applications of the invention may be made

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